

pH of Salt Solutions

Aim: To determine the pH of various salt solutions.

Equipment:

- Distilled Water
- Small beakers
- pH probe
- Universal Indicator solution
- Universal Indicator chart
- Ammonium Chloride
- Potassium Nitrate
- Sodium Acetate
- Di-sodium hydrogen phosphate
- Sodium carbonate
- Sodium chloride
- Sodium hydrogen sulfate

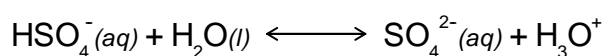
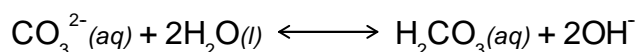
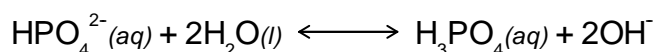
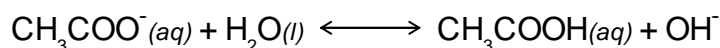
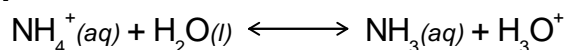
Method:

1. Place a small amount of each salt solution into separate beakers.
2. Use a pH probe to record the pH of each solution.
3. Use Universal Indicator solution and chart to determine pH of each solution.

Results:

Name of Salt	Formula	pH probe	UI pH
Ammonium Chloride	NH ₄ Cl	5.9	5
Potassium Nitrate	KNO ₃	6.4	7
Sodium Acetate	CH ₃ COONa	8.1	8
Di-sodium Hydrogen Phosphate	Na ₂ HPO ₄	8.2	8
Sodium Carbonate	Na ₂ CO ₃	9.8	10
Sodium Chloride	NaCl	6.5	7
Sodium Hydrogen Sulfate	NaHSO ₄	3.2	= 3

Calculations:



Conclusion: A successful practical.